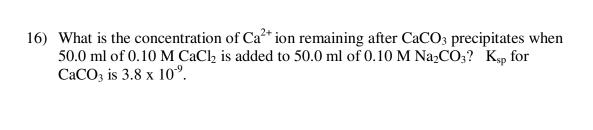
3) If a 0.05 M solution of a weak acid produces a pH of 5.9, What is K_a for this weak acid?

7) Calculate the molar solubility of barium fluoride (BaF2) in water. $K_{sp} = 1.6 \times 10^{-6}$.

Calculate the molar solubility of this compound in an aqueous solution that is 0.2 M NaF (sodium fluoride).

- 10) What is the pH of the following solutions?
 - a) 0.25 M HNO₃

b) 0.17 M Ba(OH)₂



What is the pH of a 0.15 M solution of NH₃? Kb = 1.78×10^{-5}

What is the PH of a 0.15 M solution of NH_4^+ ?