9/4/2007			CHM 191 General Chemistry - SYLLABUS 2007								
			Instructor: Dr. Chris W. Brown								
			E-mail: cbrown@chm.uri.edu								
			Office: Pastore 350 Hours: M/W/F 10-11:30, Tu/Th 8-10 am								
			Text: General Chemistry - Principles & Modern Applications								
			by Petrucci, et al, 9th Edition								
	DATE		CHAPTER								
	Sept	5	-	1	Matter - Its Properties and Measurements						
		7		1	Matter - Its Properties and Measurements						
		10		2	2 Atoms and Atomic Theory						
		12		2	Atoms and Atomic Theory						
		14		3 Chemical Compounds							
		17		3 Chemical Compounds							
		19		Exam No 1, Chap 1-3							
		21		4 Chemical Reactions							
		24		4	Chemical Reactions						
		26		5	Reactions in Aqueous Solutions						
		28		5	Reactions in Aqueous Solutions						
	Oct	1			Review						
		3		Exa	m No 2, Chap 1-5						
		5			No Class						
		8		Colu	imbus Day - no class						
		10		6	Gases (Monday classes = Lecture 9 am)						
		10		6	Gases (Monday classes = Recitation 2 pm)						
		12		6	Gases						
		15		7	Thermochemisty						
		17		7	Thermochemisty						
		19		7	Thermochemisty						
		22		7 Thermochemisty							
		24	Exam No. 3, Chap 6-7								
		26		8	Electrons in Atoms						
		29		8	Electrons in Atoms						
		31		8	Electrons in Atoms						
	Nov	2		9	Periodic Table & Some Atomic Properties						
		5		9	Periodic Table & Some Atomic Properties						
		7		10	Chemical Bonding I: Basic Concepts						
		9		10	Chemical Bonding I: Basic Concepts						
		12		Review of 8-10							
		14		Exa	m No 4, Chap 8-10						
		16		11	Chemical Bonding II: Additional Aspects						
		19		11	Chemical Bonding II: Additional Aspects						
		21		11	Chemical Bonding II: Additional Aspects						
		26		12	Liquids, Solids and Intermolcular Forces						
		28		12	Liquids, Solids and Intermolcular Forces						
		30		12	Liquids, Solids and Intermolcular Forces						
	Dec	3		Exa	m No 5, Chap 11-12						
		5		13	Solutions and Their Physical Properties						
		7		13	Solutions and Their Physical Properties						
De	ec	17 8	- 11 am		Final Exam Chapters 1 - 13						

9/4/2007		CHN	/I 191 Gei	neral Chemistry - SYLLABUS 2007					
		Instructor: Dr. Chris W. Brown (cbrown@chm.uri.edu)							
		Tea	ching As	sistant:					
		Tues	s 11 am	Drew Bordeur (dbrodeur@chm.luri.edu) Office: Pastore 250					
		Thurs 8 am		Drew Bordeur (dbrodeur@chm.luri.edu) Office: Pastore 250					
		Thurs 11 am		Drew Bordeur (dbrodeur@chm.luri.edu) Office: Pastore 250					
		Thurs 2 pm		Drew Bordeur (dbrodeur@chm.luri.edu) Office: Pastore 250					
				_					
	<u>DATE</u>		Exp #	Title					
	Sept	6		1	Scientific Measurements				
		11		1					
		13		3	Separation of Food Dyes by Paper Chromatography				
		18		3					
		20		4	Observing reactions of hosehold chemicals - deductive chemical reasoning				
		25		4					
	0.4	28		8	Determination of a chemical formula				
	UCI			8	A Cuela of Conner Depatience				
		4		5					
		9		5 27	The pH Scale and Acid Race Titrations				
		16		21					
		10		11	The Molar Volume of Dioxygen and other gases				
		23		11					
		25		14	Heat Canacity of metals				
		30		14					
	Nov	1		31	Identificationh of Alkaline Earth and Alkali Metal Ions				
		6		31					
		8		19	Molecular Models				
		13		19					
		15		18	Ionic and Covalent Bonding				
	1	20		18	Č Č				
		27		22	Colligative Properties				
		29		22					
	Dec	4		1	Final Exam ==> Written / Practical				
		6			Final Exam ==> Written / Practical				

CHM191					9/4/2007								
Course O	bjec	tives											
	То	understand the following:											
	1	Basic chemistry and chemical laws											
	2	Atoms and molecules											
	3	Chemical compounds and reactions											
	4	Solids, liquids and gases											
	5	Themochemistry - predicting products	and stabilit	es									
	6	Electrons and bonding											
	7	Periodic table and elemental properties	3										
	8	Chemical bonds											
	9	Condensed phases and their properties	S										
Requirem	ent	S:											
	1	1 Attendance in lectures, labs and recitations is required											
	2	Late assignments, problem sets and	lab report	s will not l	be accepted	d							
	3	Calculators are required for lectures	, recitation	s, labs an	d exams								
	4	Other electronic devices will not be	allowed in	exams									
Grading:													
		5 One-hour exams (drop lowest grade)	400										
		Final Exam	300										
		Quizes / Attendance / Problems	100										
		Laboratory	200										
		TOTAL	1000										
		Final grades will be based on a statistical curve. The total points (out of 1000)											
		will be plotted for all students. The top group will receive A's, the next grou											
		A-'s, the next group B+'s, etc.											
		Questions on the <i>final exam</i> will be taken from the 5 hourly exams plus question on material covered since the last (5th) hourly exam.											
		All exams will be multiple choice. The hourly exams will consist of 20											
		questions, whereas the final will have 30 questions. Exams will emphasize											
		problem solving.											
		Recitations will be held on Monday from 2 to 2:50 inPastore 234. Problems will I											
		assigned during lectures the previous week and you will be asked to work out problems on the blackboards. You will be <b>graded</b> on your ability to work											
		problems. Special Recitations - Gases and Thermochemistry: Monday, Oct 15, 2-2:50											
			Wednesday, Oct 17, 2-2:50 pm										
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