Introductory Chemistry ~ CHM 103
Course Information & Syllabus ~ Fall Semester, 2007

Instructor
Cindy Graham Brittain, PhD
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Email: cbrittain@chm.uri.edu
Office Hours: 8:30 – 10:30 AM, T – F

Lecture Materials
Custom textbook with ChemistryNow: Chemistry for Today, With Essential Study Skills (Chiras/Seager/Slabaugh)
Scientific calculator (with logarithm and exponential functions; no other electronic devices may be used in class)
Additional handouts available in class and at http://www.chm.uri.edu/chm105/index.html

Class Meetings
Section 2:   T Th  3:30 – 4:45 PM  Pastore 124
Section 3:   M W F  2 – 2:50 PM  Pastore 124

Grading Policy
The course grade will be based on the following:

- 3 Lecture Exams (24% each)  72%
- Comprehensive Final Exam  28%

Each student’s grade will be determined from the results of 4 Lecture Exams and the comprehensive Final Exam. The lowest Lecture Exam score will be dropped, and the remaining 3 exam scores will each count as 24% of the overall grade. The purpose of dropping the lowest score is to eliminate the need for make-up tests (if any students are absent on the day of a Lecture Exam). All students must adhere to the exam schedule specified on the syllabus. (The only exception to this policy will be those students who provide documentation from the University specifying the need for private testing or additional time. These arrangements must be made at least one week prior to the scheduled exam.)

Unless otherwise specified, exam questions will be primarily short answer, essay, and problem-solving. Questions will come directly from the content presented and discussed during lecture, and are likely to be similar to the recommended problems in the textbook. Students should commit to working all homework problems (those recommended in the textbook or provided as lecture handouts) as practice for taking exams; the goal is to become confident and efficient at analyzing and solving problems. Students who get the most practice solving problems tend to have the greatest success in this course. Additionally, 23 Tips for “Improving Your Test-taking Abilities” are provided in Chiras Chapter 3.

If the Kingston campus is closed due to weather (or other unexpected event) on a scheduled Lecture Exam day, students should expect that the exam will be given at the next class meeting.

Students should anticipate that graded exam papers will be returned one week after the exam date (via the Pastore hallway mailboxes).

Lecture Attendance, Preparation, and Participation
Students need to plan time for significant study outside of class. A rule of thumb for most chemistry courses: 2 to 3 hours of study time outside of class for every 1 hour of in-class lecture. As lecture meets 2 ½ hours each week, this means that students should plan to study 5 to 7 ½ hours every week.

Students should prepare to participate actively in each lecture by:

- Becoming familiar with the topic of each lecture ahead of time
- Spending time before each lecture reviewing previous notes
- Practicing the recommended problems related to current lecture topics

Additionally, it’s recommended that students closely adhere to the 19 Tips for “Getting the Most out of Lectures and Labs” in Chiras Chapter 1, and the 17 Tips for “Getting the Most Out of Reading Assignments” in Chiras Chapter 2.
CONSISTENT LECTURE ATTENDANCE IS ESSENTIAL FOR SUCCESS IN THIS COURSE. The most significant mistake a student can make in this course is to miss lecture. The 2nd most significant mistake is to attend lecture, but be inattentive (or worse, disruptive, and thus distracting to those seated nearby).

Each student should identify several study partners that are in the same lecture section. In the event of an absence, students should plan to immediately obtain copies of missed lecture content, hand-outs, and announcements from their study partners.

STUDENTS CANNOT AFFORD TO FALL BEHIND IN THIS COURSE. Every new concept builds on material that students should have previously mastered. Students must take responsibility for reviewing previous material as needed.

Study Help Resources

- Chemistry Graduate Student Teaching Assistants in the Chemistry TA Help Office, Pastore 215. Note that not all TAs can provide help with all courses, and students seeking help should refer to the schedule of office hours posted at the door.

- Tutors at the Academic Enhancement Center (AEC). The AEC in Roosevelt Hall offers a comfortable environment in which to study alone or in groups, with or without a tutor. AEC tutors are fellow students who have taken the courses, and they can answer questions, clarify concepts, check for understanding, and help with study. Students are encouraged to make an appointment, but are also welcome to simply walk in during office hours: M – Th (9 AM – 9 PM), F (9 AM – 1 PM), and Sun (4 PM – 8 PM). To obtain a complete schedule (including when tutors are available specifically for this class), students can refer to the AEC websites at [www.uri.edu/aec](http://www.uri.edu/aec), call (401) 874-2367, or stop by the fourth floor in Roosevelt Hall.

Academic Honesty

Academic dishonesty in any form is considered a serious offense, and disciplinary action will be taken immediately. The university’s policy on academic honesty is detailed in the student handbook (available online), and it is summarized below:

Students are expected to be honest in all academic work. A student’s name on any written work, including assignments, lab reports, internship reports, papers, or examinations, shall be regarded as assurance that the work is the result of the student’s own thought and study. Work should be stated in the student’s own words, and produced without assistance or properly attributed to its source. When students are authorized to work jointly, group effort must be indicated on the work submitted.

The following are examples of academic dishonesty:

- Unauthorized communication during exams.
- Unauthorized use of another’s work or preparing work for another student.
- Taking an exam for another student.
- Altering or attempting to alter grades.
- The use of notes or electronic devices such as calculators, computers, or cell phones to gain an unauthorized advantage during exams.
- Fabricating or falsifying facts, data, or references.
- Facilitating or aiding another’s academic dishonesty.

When there is an allegation of academic dishonesty, the instructor may:

- Fail the student for the assignment.
- Recommend that the student fail the course.
### Syllabus – Section 2

<table>
<thead>
<tr>
<th>Week #</th>
<th><strong>TUESDAY</strong></th>
<th><strong>THURSDAY</strong></th>
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</thead>
</table>
| 1      | 9/4: No Classes – Advising Day | 9/6: Course Information  
* S/S Ch 1: Matter, Measurements, and Calculations |
| 2      | 9/11:  
* S/S Ch 1: Matter, Measurements, and Calculations | 9/13:  
* S/S Ch 1: Matter, Measurements, and Calculations |
| 3      | 9/18:  
* S/S Ch 2: Atoms and Molecules | 9/20:  
* S/S Ch 2: Atoms and Molecules |
| 4      | 9/25: EXAM 1 | 9/27:  
* S/S Ch 3: Electronic Structure and the Periodic Law |
| 5      | 10/2:  
* S/S Ch 3: Electronic Structure and the Periodic Law | 10/4:  
* S/S Ch 4: Forces between Particles |
| 6      | 10/9:  
* S/S Ch 4: Forces between Particles | 10/11:  
* S/S Ch 4: Forces between Particles |
| 7      | 10/16: EXAM 2 | 10/18:  
* S/S Ch 5: Chemical Reactions |
| 8      | 10/23:  
* S/S Ch 5: Chemical Reactions | 10/25:  
* S/S Ch 6: The States of Matter |
| 9      | 10/30:  
* S/S Ch 6: The States of Matter | 11/1:  
* S/S Ch 7: Solutions and Colloids |
| 10     | 11/6:  
* S/S Ch 7: Solutions and Colloids | 11/8:  
* S/S Ch 8: Reaction Rates and Equilibria |
| 11     | 11/13:  
* S/S Ch 8: Reaction Rates and Equilibria | 11/15: EXAM 3 |
| 12     | 11/20:  
* S/S Ch 8: Reaction Rates and Equilibria | 11/22:  
* No Classes – Thanksgiving Day Holiday |
| 13     | 11/27:  
* S/S Ch 9: Acids, Bases, and Salts | 11/29:  
* S/S Ch 9: Acids, Bases, and Salts |
| 14     | 12/4:  
* S/S Ch 9: Acids, Bases, and Salts | 12/6: EXAM 4 |
| 15     | 12/11: No Classes – Reading Day  
Optional Help Session to prepare for Final Exam |  |

**FINAL EXAM:** Tuesday, December 18th, 3:00 – 6:00 PM.
Syllabus – Section 3

<table>
<thead>
<tr>
<th>Week #</th>
<th>MONDAY</th>
<th>WEDNESDAY</th>
<th>FRIDAY</th>
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</thead>
<tbody>
<tr>
<td>1</td>
<td>9/3: No Classes – Labor Day Holiday</td>
<td>9/5: Course Information&lt;br&gt; S/S Ch 1: Matter, Measurements, and Calculations</td>
<td>9/6: S/S Ch 1: Matter, Measurements, and Calculations</td>
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<td>5</td>
<td>10/1: S/S Ch 3: Electronic Structure and the Periodic Law</td>
<td>10/3: S/S Ch 4: Forces between Particles</td>
<td>10/5: S/S Ch 4: Forces between Particles</td>
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<tr>
<td>6</td>
<td>10/8: No Classes – Columbus Day Holiday</td>
<td>10/10: S/S Ch 4: Forces between Particles</td>
<td>10/12: S/S Ch 4: Forces between Particles</td>
</tr>
<tr>
<td>7</td>
<td>10/15: S/S Ch 4: Forces between Particles</td>
<td>10/17: EXAM 2</td>
<td>10/19: S/S Ch 5: Chemical Reactions</td>
</tr>
<tr>
<td>10</td>
<td>11/5: S/S Ch 7: Solutions and Colloids</td>
<td>11/7: S/S Ch 7: Solutions and Colloids</td>
<td>11/9: S/S Ch 8: Reaction Rates and Equilibria</td>
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<td>11</td>
<td>11/12: No Classes – Veterans’ Day Holiday</td>
<td>11/14: S/S Ch 8: Reaction Rates and Equilibria</td>
<td>11/16: EXAM 3</td>
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<tr>
<td>15</td>
<td>12/10: Last Day of Classes&lt;br&gt;Lecture/Review</td>
<td>12/12: Reading Day – Optional Help&lt;br&gt;Session to prepare for Final Exam</td>
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FINAL EXAM: Wednesday, December 19th, 8:00 – 11:00 AM.